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1023 mol⁻¹ Specific heat capacity of water, $c = 4.18 \text{ J g}^{-1} \text{ K}^{-1}$ Ionic product of water, $K_w = 1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ at 298 K 1 tonne = 106 g Arrhenius equation: $k = Ae^{-E_a/RT}$ or $\ln k = -E_a/RT + \ln A$... [Ptable.com Periodic Table](#)
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